



## **Formation and function of molecules depending on chemical bonding**

### **[ GLOSSARY ]**

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# **GLOSSARY**

### **Atoms**

The smallest particle of a chemical element that can exist or basic unit of element

### **Dimer**

a molecule or molecular complex consisting of two identical molecules linked together.

### **Dipole**

A dipole is a separation of electrical charges. In chemistry, a dipole refers to the separation of charges within a molecule between two covalently bonded atoms.

### **Electronegativity**

It is a measure of the tendency of an atom to attract a bonding pair of electrons.

### **Electrostatic force**

It is also known as Coulomb force. It is an interaction between objects or particles due to their electric charges.

### **Molecule**

A group of atoms bonded together, representing the smallest fundamental unit of a chemical compound that can take part in a chemical reaction.

**Noble gas**

Any of the gaseous elements helium, neon, argon, krypton, xenon, and radon occupying Group 0 (18) of the periodic table.

**Octet (rule)**

The octet rule is a chemical rule of thumb that states that atoms of main-group elements tend to combine in such a way that each atom has eight electrons in its valence shell, giving it the same electronic configuration as a noble gas.

**Polar bond**

It is a type of covalent bond between two atoms in which electrons are shared unequally. Because of this, one end of the molecule has a slightly negative charge and the other a slightly positive charge.

**Valence shell**

It is the outermost shell of an atom containing the valence electrons.